**Shapes of molecules (Level 2) examiners tips: Read these please!**

• use molypods, plastercine & toothpicks or balloons etc to help you to visualise the shapes in 3D

• make and rotate interactive animations of molecules

• use [PHET molecule basics](https://phet.colorado.edu/en/simulation/molecule-shapes-basics), [molecule3d](http://chemicalminds.wikispaces.com/file/view/molecule3d.swf) and [pbstutorial](http://www.pbslearningmedia.org/asset/lsps07_int_molecularshp/) to reinforce your visual images of the different shapes

• recognise and state the number of regions (areas) of electron repulsion around a central atom

• learn the names of the shapes and their angles

• learn them, no excuses!

• you are determining the shape around the central atom

• no, you aren’t required to draw the lone pairs or valence electrons when drawing SHAPES of molecules

• yes, an alternative accepted word for trigonal is triangular

• see below a [.pdf summary](l2shapessummary.pdf) of the shapes of molecules for Level 2, also click below to download as a [Powerpoint](summaryl2shapes.pptx)

**Also…”don’t be daft”**

a single or double or triple bond are all ONE REGION of negative charge

the trigonal planar shape is VERY, VERY different to trigonal pyramidal shape

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